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## Electrons in Atoms pt 1

Chapter 9 - Electrons in atoms and the Periodic Table  
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Although the speed of all electromagnetic waves in a vacuum is the same, waves can have different wavelengths and frequencies. As you can see from the equation on the previous page, wavelength and frequency are inversely related; in other words, as one quantity increases, the other decreases.

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atom is a tiny, indestructible particle with no internal structure. Thomson. The atom is a sphere of positive electrical charge with electrons embedded in the sphere.

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Duality/Electromagnetic  
Spectrum/Relationship of  
Wavelength, Frequency and Speed of  
light 5.2 Bohr's Model of the  
Atom/Quantum Mechanical Model of  
the Atom 5.3 Electron Arrangement &  
Valence Electrons.

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Section 5.2 – Electron Arrangement  
in Atoms The electron configuration  
of an atom is the arrangement of the  
electrons. There are 3 rules that  
govern the electron configuration:  
Aufbau ' s principle, Pauli Exclusion  
principle, and Hund ' s rule.

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Chapter 5 “ Electrons in Atoms ”

Chemistry Charles Page High School

Stephen L. Cotton \* \* \* \* \* The

electromagnetic spectrum consists of radiation over a broad band of wavelengths. The visible light portion is very small. It is in the 10-7m wavelength range and 10<sup>15</sup> Hz (s<sup>-1</sup>) frequency range.

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Unit 4: Electrons in Atoms (Chapter 5)

You'll Remember ...

Chapter 5 Electrons in Atoms. STUDY. PLAY. Quantum Mechanical Model. model of the atom we believe today that involves the probability of finding an electron in a certain position. What is the maximum number of f orbitals in any single energy level in an atom ?  
7. Electrons in the same orbital.

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CHAPTER 5 What You 'll Learn You will compare the wave and particle models of light. You will describe how the frequency of light emitted by an atom is a unique characteristic of that atom. You will compare and contrast the Bohr and quantum mechanical

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138 Chapter 5 Electrons in Atoms  
Electron Configurations for Elements  
in Period Three Table 5-4 Figure 5-19.  
This sublevel diagram shows the  
order in which the orbitals are usually  
filled. The proper sequence for the  
first seven orbitals is 1s, 2s, 2p, 3s,  
3p, 4s, and 3d. Chapter 5 Electrons in  
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of the Atom • Rutherford used existing ideas about the atom and proposed an atomic model in which the electrons move around the nucleus, like the planets move around the sun.

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Rutherford ' s model fails to explain why objects change color when heated.

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CHAPTER 5 What You ' ll Learn You will compare the wave and particle models of light. You will describe how the frequency of light emitted by an atom is a unique characteristic of that atom. You will compare and contrast the Bohr and quantum mechanical models of the atom. You will express the arrangements of electrons in

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How many electrons can each p orbital hold? Chapter 5: Electrons in Atoms DRAFT. 10th - 11th grade. 60 times. Chemistry. 77% average accuracy. 2 years ago. msrlyounger. 0. Save. Edit. Edit. Chapter 5: Electrons in Atoms DRAFT. 2 years ago. by msrlyounger. Played 60 times. 0. 10th - 11th grade .

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Chapter 5 Electrons in Atoms 2. Light and Quantized Energy (5.1)

- The study of light led to the development of the quantum mechanical model.

- Light is a kind of

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electromagnetic radiation (EM).

All move at  $3.00 \times 10^8$  m/s (c) Speed of light.

3.

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